

## 4-Bromo-2,6-dimethylanilinium bromide monohydrate

Jun-jie Tang,<sup>a</sup> Jing Chen,<sup>a\*</sup> Jin-tang Wang,<sup>a</sup> Ai-hua Lu<sup>b</sup> and You-sheng Chen<sup>c</sup>

<sup>a</sup>Department of Applied Chemistry, College of Science, Nanjing University of Technology, Xinmofan Road No. 5 Nanjing, Nanjing 210009, People's Republic of China, <sup>b</sup>Nanjing Chemzham Pharmaceutical Technology Co. Ltd, Nanjing 210017, People's Republic of China, and <sup>c</sup>Department of Pharmacy, Guangdong Vocational and Technical College of Chemical Engineering Pharmaceutics, Guangzhou 510520, People's Republic of China

Correspondence e-mail: wjt@njut.edu.cn

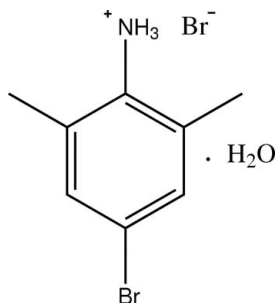
Received 15 November 2007; accepted 1 December 2007

Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.008$  Å;  $R$  factor = 0.045;  $wR$  factor = 0.096; data-to-parameter ratio = 17.3.

In the title compound,  $\text{C}_8\text{H}_{11}\text{BrN}^+\cdot\text{Br}^-\cdot\text{H}_2\text{O}$ , a network of  $\text{N}-\text{H}\cdots\text{O}$ ,  $\text{N}-\text{H}\cdots\text{Br}$  and  $\text{O}-\text{H}\cdots\text{Br}$  hydrogen bonds helps to consolidate the crystal packing.

### Related literature

For related literature, see: Hirao & Fukuhara (1998); MacDiamid *et al.* (1998); Wakayama (1998); Wang *et al.* (2002).



### Experimental

#### Crystal data

$\text{C}_8\text{H}_{11}\text{BrN}^+\cdot\text{Br}^-\cdot\text{H}_2\text{O}$   
 $M_r = 298.99$   
 Monoclinic,  $P2_1/n$

$a = 7.1630$  (14) Å  
 $b = 18.649$  (4) Å  
 $c = 8.4770$  (17) Å

$\beta = 109.98$  (3)°  
 $V = 1064.2$  (4) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation

$\mu = 7.58$  mm<sup>-1</sup>  
 $T = 293$  (2) K  
 $0.40 \times 0.20 \times 0.20$  mm

#### Data collection

Enraf–Nonius CAD-4 diffractometer  
 Absorption correction:  $\psi$  scan (North *et al.*, 1968)  
 $T_{\min} = 0.152$ ,  $T_{\max} = 0.221$   
 2077 measured reflections

2077 independent reflections  
 1346 reflections with  $I > 2\sigma(I)$   
 3 standard reflections every 200 reflections  
 intensity decay: none

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.045$   
 $wR(F^2) = 0.096$   
 $S = 0.97$   
 2077 reflections  
 120 parameters  
 2 restraints

H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\max} = 0.42$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.52$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{O1}-\text{H2}\cdots\text{Br2}^{\text{i}}$	0.84 (2)	2.56 (3)	3.380 (5)	169 (3)
$\text{O1}-\text{H1}\cdots\text{Br2}$	0.84 (2)	2.50 (2)	3.322 (5)	174 (3)
$\text{N1}-\text{H1C}\cdots\text{O1}$	0.89	1.87	2.754 (6)	170
$\text{N1}-\text{H1B}\cdots\text{Br2}^{\text{ii}}$	0.89	2.53	3.388 (5)	163
$\text{N1}-\text{H1A}\cdots\text{Br2}^{\text{iii}}$	0.89	2.71	3.523 (4)	153

Symmetry codes: (i)  $x - \frac{1}{2}, -y + \frac{3}{2}, z - \frac{1}{2}$ ; (ii)  $x + \frac{1}{2}, -y + \frac{3}{2}, z - \frac{1}{2}$ ; (iii)  $x, y, z - 1$ .

Data collection: *CAD-4 Software* (Enraf–Nonius, 1989); cell refinement: *CAD-4 Software*; data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *SHELXL97*; software used to prepare material for publication: *SHELXL97*.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: LW2046).

### References

- Enraf–Nonius (1989). *CAD-4 Software*. Version 5.0. Enraf–Nonius, Delft, The Netherlands.
- Harms, K. & Wocadlo, S. (1995). *XCAD4*. University of Marburg, Germany.
- Hirao, T. & Fukuhara, S. (1998). *J. Org. Chem.* **63**, 7534–7535.
- MacDiamid, A. G., Zhang, W. J., Feng, J., Huang, F. & Hsieh, B. (1998). *Polym. Prepr. (Am. Chem. Soc. Div. Polym. Chem.)*, **39**, 80–81.
- North, A. C. T., Phillips, D. C. & Mathews, F. S. (1968). *Acta Cryst.* **A24**, 351–359.
- Sheldrick, G. M. (1997). *SHELXL97* and *SHELXS97*. University of Göttingen, Germany.
- Wakayama, D. K. K. (1998). *Shizen Kagaku*, **48**, 9–15.
- Wang, C., Gao, J. B. & Chen, C. H. (2002). *Polym. Prepr. (Am. Chem. Soc. Div. Polym. Chem.)*, **40**, 1746–1747.

**supplementary materials**

*Acta Cryst.* (2008). E64, o244 [ doi:10.1107/S1600536807064896 ]

## 4-Bromo-2,6-dimethylanilinium bromide monohydrate

J. Tang, J. Chen, J. Wang, A. Lu and Y. Chen

### Comment

aniline is a novel, useful intermediate, which can be used in various areas. Some derivatives of aniline have improving anti-corrosion ability for metals (Wang *et al.*, 2002). Some show high efficacy as chemical sensors (MacDiamid *et al.*, 1998) and catalytic oxidation (Hirao & Fukuhara, 1998). We report here the crystal structure of the title compound, (I). A network of intermolecular N—H···Br and O—H···Br hydrogen bonds helps to establish the crystal packing, Fig. 1 and Fig. 2.

### Experimental

The title compound is synthesized according to the literature (Wakayama *et al.*, 1998). Crystals of (I) suitable for X-ray diffraction were obtained by slow evaporation of an ethanol solution.

### Refinement

All H atoms bonded to the C atoms were placed geometrically at the distances of 0.93–0.96 Å and included in the refinement in riding motion approximation with  $U_{\text{iso}}(\text{H}) = 1.2$  or  $1.5U_{\text{eq}}$  of the carrier atom.

### Figures

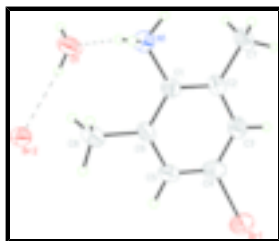


Fig. 1. A view of the molecular structure of (I). Displacement ellipsoids are drawn at the 30% probability level. The crystal structure of (I).

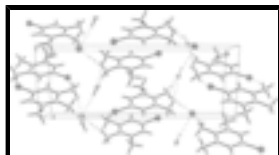


Fig. 2. A packing diagram of the title molecular structure, with dashed lines indicate intermolecular N—H···O, N—H···Br and O—H···Br hydrogen bonds.

## 4-Bromo-2,6-dimethylanilinium bromide monohydrate

### Crystal data

$\text{C}_8\text{H}_{11}\text{BrN}^+\cdot\text{Br}^-\cdot\text{H}_2\text{O}$

$M_r = 298.99$

Monoclinic,  $P2_1/n$

$F_{000} = 584$

$D_x = 1.866 \text{ Mg m}^{-3}$

Mo  $K\alpha$  radiation

# supplementary materials

---

Hall symbol: -P 2yn

$a = 7.1630 (14) \text{ \AA}$

$b = 18.649 (4) \text{ \AA}$

$c = 8.4770 (17) \text{ \AA}$

$\beta = 109.98 (3)^\circ$

$V = 1064.2 (4) \text{ \AA}^3$

$Z = 4$

$\lambda = 0.71073 \text{ \AA}$

Cell parameters from 25 reflections

$\theta = 9\text{--}13^\circ$

$\mu = 7.58 \text{ mm}^{-1}$

$T = 293 (2) \text{ K}$

Block, colorless

$0.40 \times 0.20 \times 0.20 \text{ mm}$

## Data collection

Enraf–Nonius CAD-4  
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 293(2) \text{ K}$

$\omega/2\theta$  scans

Absorption correction:  $\psi$  scan  
(North *et al.*, 1968)

$T_{\min} = 0.152$ ,  $T_{\max} = 0.221$

2077 measured reflections

2077 independent reflections

1346 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.0000$

$\theta_{\max} = 26.0^\circ$

$\theta_{\min} = 2.2^\circ$

$h = -8 \rightarrow 8$

$k = 0 \rightarrow 22$

$l = 0 \rightarrow 10$

3 standard reflections

every 200 reflections

intensity decay: none

## Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.045$

$wR(F^2) = 0.096$

$S = 0.97$

2077 reflections

120 parameters

2 restraints

Primary atom site location: structure-invariant direct  
methods

Secondary atom site location: difference Fourier map

Hydrogen site location: difference Fourier map

H atoms treated by a mixture of  
independent and constrained refinement

$w = 1/[\sigma^2(F_o^2) + (0.0437P)^2]$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.42 \text{ e \AA}^{-3}$

$\Delta\rho_{\min} = -0.52 \text{ e \AA}^{-3}$

Extinction correction: none

## Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > 2\sigma(F^2)$  is used only for calculat-

ing  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ - factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
Br1	0.93285 (11)	0.38204 (3)	0.77315 (8)	0.0490 (2)
Br2	0.47862 (10)	0.69675 (3)	0.95867 (8)	0.0407 (2)
C1	0.7335 (8)	0.5937 (3)	0.4804 (7)	0.0276 (13)
C2	0.6858 (8)	0.5297 (3)	0.3931 (7)	0.0310 (13)
C3	0.7483 (8)	0.4664 (3)	0.4828 (7)	0.0346 (14)
H3	0.7195	0.4224	0.4282	0.041*
C4	0.8528 (9)	0.4691 (3)	0.6521 (7)	0.0368 (15)
C5	0.9007 (8)	0.5334 (3)	0.7373 (7)	0.0330 (14)
H5	0.9728	0.5338	0.8517	0.040*
C6	0.8412 (8)	0.5966 (3)	0.6519 (6)	0.0298 (13)
C7	0.5756 (9)	0.5250 (3)	0.2060 (6)	0.0410 (16)
H7A	0.4469	0.5466	0.1799	0.062*
H7B	0.5605	0.4755	0.1723	0.062*
H7C	0.6494	0.5497	0.1472	0.062*
C8	0.8945 (9)	0.6665 (3)	0.7470 (7)	0.0420 (16)
H8A	0.7752	0.6910	0.7436	0.063*
H8B	0.9682	0.6960	0.6964	0.063*
H8C	0.9739	0.6569	0.8616	0.063*
N1	0.6621 (7)	0.6625 (2)	0.3935 (5)	0.0336 (12)
H1A	0.6039	0.6545	0.2841	0.050*
H1B	0.7645	0.6921	0.4099	0.050*
H1C	0.5749	0.6822	0.4343	0.050*
O1	0.4192 (8)	0.7175 (3)	0.5552 (6)	0.0494 (12)
H1	0.442 (3)	0.710 (4)	0.657 (3)	0.05 (2)*
H2	0.306 (3)	0.736 (4)	0.518 (3)	0.09 (3)*

*Atomic displacement parameters ( $\text{\AA}^2$ )*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Br1	0.0650 (5)	0.0365 (4)	0.0427 (4)	0.0084 (4)	0.0147 (3)	0.0129 (3)
Br2	0.0445 (4)	0.0422 (4)	0.0337 (3)	0.0045 (3)	0.0113 (3)	0.0053 (3)
C1	0.030 (3)	0.028 (3)	0.027 (3)	-0.001 (2)	0.012 (3)	0.000 (2)
C2	0.035 (3)	0.033 (3)	0.027 (3)	-0.001 (3)	0.014 (3)	-0.006 (3)
C3	0.040 (4)	0.027 (3)	0.034 (3)	-0.005 (3)	0.011 (3)	-0.005 (3)
C4	0.038 (4)	0.039 (4)	0.033 (3)	-0.001 (3)	0.011 (3)	0.010 (3)
C5	0.037 (4)	0.034 (3)	0.024 (3)	0.002 (3)	0.006 (3)	0.002 (3)
C6	0.031 (3)	0.034 (3)	0.023 (3)	-0.001 (3)	0.008 (3)	-0.006 (2)
C7	0.054 (4)	0.038 (4)	0.025 (3)	-0.006 (3)	0.006 (3)	-0.009 (3)
C8	0.045 (4)	0.048 (4)	0.027 (3)	0.003 (3)	0.003 (3)	-0.004 (3)
N1	0.037 (3)	0.032 (3)	0.033 (3)	0.002 (2)	0.012 (2)	0.000 (2)
O1	0.052 (3)	0.069 (4)	0.026 (3)	0.019 (3)	0.012 (2)	0.005 (2)

## supplementary materials

---

### Geometric parameters (Å, °)

Br1—C4	1.901 (6)	C7—H7A	0.9600
C1—C2	1.384 (7)	C7—H7B	0.9600
C1—C6	1.395 (7)	C7—H7C	0.9600
C1—N1	1.481 (6)	C8—H8A	0.9600
C2—C3	1.393 (7)	C8—H8B	0.9600
C2—C7	1.513 (7)	C8—H8C	0.9600
C3—C4	1.374 (7)	N1—H1A	0.8900
C3—H3	0.9300	N1—H1B	0.8900
C4—C5	1.381 (8)	N1—H1C	0.8900
C5—C6	1.373 (7)	O1—H1	0.84 (2)
C5—H5	0.9300	O1—H2	0.84 (2)
C6—C8	1.511 (7)		
C2—C1—C6	122.6 (5)	C2—C7—H7B	109.5
C2—C1—N1	120.0 (5)	H7A—C7—H7B	109.5
C6—C1—N1	117.3 (5)	C2—C7—H7C	109.5
C1—C2—C3	117.6 (5)	H7A—C7—H7C	109.5
C1—C2—C7	123.8 (5)	H7B—C7—H7C	109.5
C3—C2—C7	118.5 (5)	C6—C8—H8A	109.5
C4—C3—C2	119.9 (5)	C6—C8—H8B	109.5
C4—C3—H3	120.1	H8A—C8—H8B	109.5
C2—C3—H3	120.1	C6—C8—H8C	109.5
C3—C4—C5	121.8 (5)	H8A—C8—H8C	109.5
C3—C4—Br1	119.3 (5)	H8B—C8—H8C	109.5
C5—C4—Br1	118.9 (4)	C1—N1—H1A	109.5
C6—C5—C4	119.5 (5)	C1—N1—H1B	109.5
C6—C5—H5	120.2	H1A—N1—H1B	109.5
C4—C5—H5	120.2	C1—N1—H1C	109.5
C5—C6—C1	118.5 (5)	H1A—N1—H1C	109.5
C5—C6—C8	118.9 (5)	H1B—N1—H1C	109.5
C1—C6—C8	122.6 (5)	H1—O1—H2	106 (3)
C2—C7—H7A	109.5		
C6—C1—C2—C3	0.4 (8)	C3—C4—C5—C6	0.7 (9)
N1—C1—C2—C3	-176.9 (5)	Br1—C4—C5—C6	-179.0 (4)
C6—C1—C2—C7	-177.7 (5)	C4—C5—C6—C1	0.0 (8)
N1—C1—C2—C7	5.0 (8)	C4—C5—C6—C8	179.9 (5)
C1—C2—C3—C4	0.2 (9)	C2—C1—C6—C5	-0.6 (8)
C7—C2—C3—C4	178.5 (5)	N1—C1—C6—C5	176.8 (5)
C2—C3—C4—C5	-0.8 (9)	C2—C1—C6—C8	179.5 (6)
C2—C3—C4—Br1	178.9 (4)	N1—C1—C6—C8	-3.1 (8)

### Hydrogen-bond geometry (Å, °)

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
O1—H2 $\cdots$ Br2 <sup>i</sup>	0.84 (2)	2.56 (3)	3.380 (5)	169 (3)
O1—H1 $\cdots$ Br2	0.84 (2)	2.50 (2)	3.322 (5)	174 (3)
N1—H1C $\cdots$ O1	0.89	1.87	2.754 (6)	170

N1—H1B···Br2 <sup>ii</sup>	0.89	2.53	3.388 (5)	163
N1—H1A···Br2 <sup>iii</sup>	0.89	2.71	3.523 (4)	153

Symmetry codes: (i)  $x-1/2, -y+3/2, z-1/2$ ; (ii)  $x+1/2, -y+3/2, z-1/2$ ; (iii)  $x, y, z-1$ .

Fig. 1

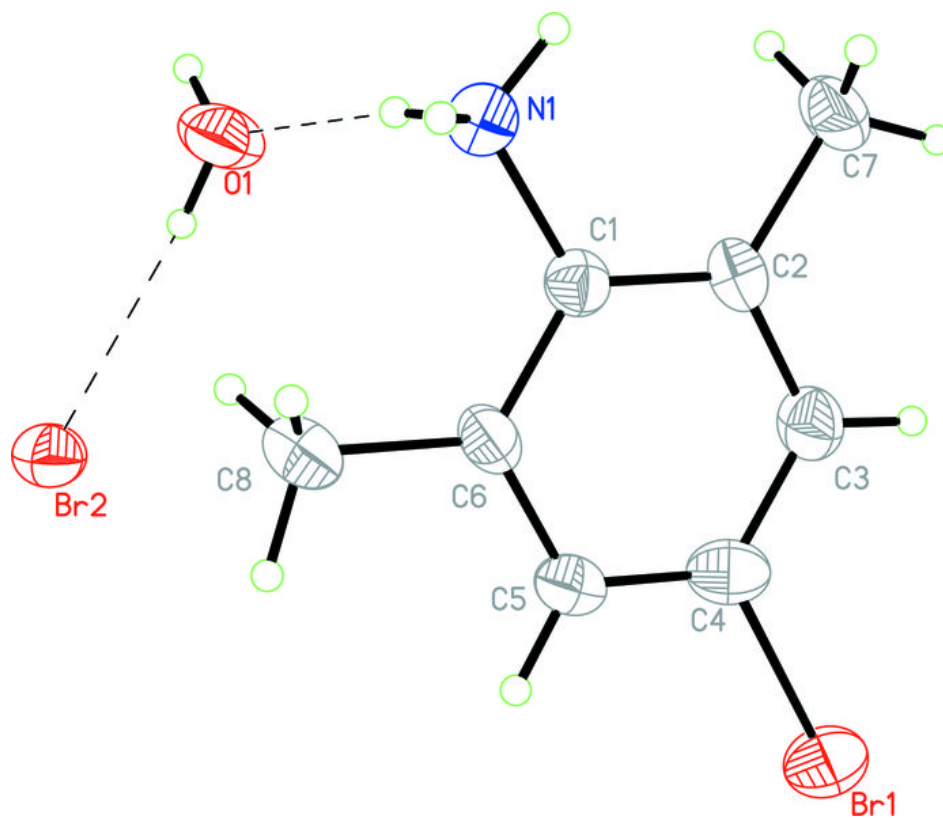




Fig. 2

